

Charge-Transfer Intermediates in the Electrochemical Doping Mechanism of Conjugated Polymers

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Abstract

We address the nature of electrochemically induced charged states in conjugated polymers, their evolution as a function of electrochemical potential, and their coupling to their local environment by means of transient absorption and Raman spectroscopies synergistically performed *in situ* throughout the electrochemical doping process. In particular, we investigate the fundamental mechanism of electrochemical doping in an oligoether-functionalized 3,4-propylenedioxythiophene (ProDOT) copolymer. The changes embedded in both linear and

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transient absorption features allow us to identify a precursor electronic state with charge-transfer (CT) character that precedes polaron formation and bulk electronic conductivity. This state is shown to contribute to the ultrafast quenching of both neutral molecular excitations and polarons. Raman spectra relate the electronic transition of this precursor state predominantly to the C_{β} - C_{β} stretching mode of the thiophene heterocycle. We characterize the coupling of the CT-like state with primary excitons and electrochemically induced charge-separated states, providing insight into the energetic landscape of a heterogeneous polymer-electrolyte system and demonstrating how such coupling depends on environmental parameters, such as polymer structure, electrolyte composition, and environmental polarity.

Introduction

Organic mixed ionic-electronic conductors (OMIECs) constitute a broad family of mostly π -conjugated polymers with the remarkable ability to support facile and simultaneous transport of ionic and electronic charged species. During electrochemical doping, charge carriers are formed along the conjugated polymer backbone in response to a sufficiently large change in potential of the underlying electrode to initiate a charge transfer across the OMIEC-electrode interface. Simultaneously, solvated electrolyte counter-ions move across the OMIEC-electrolyte interface and into the bulk of the active OMIEC layer to stabilize the induced charge carriers. Accordingly, the transport and coupling of ionic and electronic species in OMIECs involves interactions and processes spanning multiple length scales and interfaces. As such, mixed transport in conjugated polymers is influenced by an assortment of factors, including the choice of polymer structure, the extent of inter- and intrachain electronic coupling, the mesoscale assembly of polymer chains in the active layer, the ability of the polymer to interact with its ionic environment, and the nature of the polymer-electrode interface.

Future technological development relies on understanding both ion and electron transport individually along with the coupling of these two properties and the elementary processes that govern charge formation and stabilization.¹⁻³ Such information is relevant across a wide range of applications that rely on OMIECs, such as electrochemical charge storage devices (batteries and supercapacitors),⁴⁻⁸ electrochromic displays,⁹⁻¹¹ light-emitting electrochemical cells,^{12,13} as well as several emerging bioelectronic technologies,^{3,14} including organic electrochemical transistors

(OECTs),^{15–19} ion pumps,^{20,21} sensors,^{22,23} and soft actuators.^{24,25} Regardless of the application, all of these devices involve polymeric active layers that undergo continuous changes in their oxidation state (doping level) due to local variations in electrochemical potential. Such doping processes result in significant, potential-dependent fluctuations in polymer microstructure, energetic landscape, electrolyte solvation, ion content, and electronic density of states. An additional layer of complexity is introduced by the weak intermolecular coupling between polymer chains and numerous degrees of conformational freedom in organic materials. Combined with the high charge density required for electrochemical doping, this often yields heterogeneous potential energy landscapes in the OMIEC, where different regions or domains of a single active layer can vary greatly in terms of their electroactivity and charge transport properties. In some cases, heterogeneity occurs over extremely small length scales comprising only a few chromophore segments, which greatly limits the utility of conventional characterization techniques that rely on longer-range order and/or periodicity (e.g. X-ray and scanning probe methods^{26–30}) for understanding the doping process. It is therefore necessary to develop new approaches for capturing the dynamic picture underlying electrochemical doping to provide a molecular description of charged states in weakly-coupled, structurally irregular, and heterogeneous polymer-electrolyte systems. Specifically, a molecular-level understanding of the nature of all electronic species involved in the electrochemical doping process, including electronic intermediates and their relationship to fully de-doped and fully doped states, is required for the rational design of novel OMIECs with the greatest performance in wide-ranging applications.

To address part of this growing challenge, we apply a unique combination of both linear and ultrafast spectroscopies *in situ* to probe the nature and dynamics of the primary photoexcitations and the ground-state polaron population in a model OMIEC while modulating its electrochemical doping level. Investigating the excited-state dynamics of photoexcitations in this manner allows us to gain insight into the formation and evolution of charge carriers, the coupling of such excitations to their environment (i.e. molecular vibrations, electronic fluctuations, electrolyte composition, other excited states, etc.), and the interaction of these states with neighboring excitons and polarons – all as a function of the electrochemical potential.

Here, we characterize the dynamics of primary photoexcitations in an oligoether-functionalized 3,4-propylenedioxythiophene copolymer (ProDOT(OE₃)-DMP) and evaluate the

correlation of these excitations with the evolution of macromolecular and environmental properties as the doping process unfolds. As previously reported, ProDOT(OE₃)-DMP can be fully and reversibly p-doped in both aqueous and organic electrolytes due to its electronic-rich and sterically protected backbone, low oxidation potential, and polar side chains,³¹ making this material uniquely suited for these studies. Incorporation of oligoether groups, which facilitate electrochemical doping in aqueous media, is a common structural approach for designing high performance OMIECs, especially in devices that require electrochemical interfacing with biological systems.^{32–36} By performing measurements *in situ* with precise control of the polymer's doping level, we develop an understanding regarding the elementary formation of charge carriers and their coupling to their local environment in a model p-type OMIEC.

In particular, we present results demonstrating that the electrochemical doping of ProDOT(OE₃)-DMP is mediated by a sub-gap state which emerges prior to the formation of mobile charge carriers. This state exhibits properties consistent with a polarized molecular excitation, i.e. a fundamental electronic excitation that has acquired a charge-transfer (CT) character, in which the electron and hole density is weighted unequally in different parts of the chromophore. Assignment of this species is supported by resonance Raman, linear absorption, and transient absorption spectroscopies performed incrementally throughout the electrochemical doping process. Together, these techniques show that this CT-like state (*i*) is present prior to the onset of redox current and is highly sensitive to the chemical composition of the surrounding environment, (*ii*) exhibits a potential dependence distinct from the traditionally assigned “polaron” feature,³⁷ (*iii*) shares a common ground state with the exciton of the undoped polymer, and (*iv*) plays a substantial role in the ultrafast quenching of the excitonic population. Depending on the environmental parameters (protic and aprotic systems, ion size and structure, polarity, hydrogen bonding, pH, etc.), this state could play an active role in the doping mechanism by favoring the reorganization of the electronic configuration from the neutral to the doped form: understanding its nature, coupling to the environment, and relationship to other electronic species is therefore a necessary step for an informed design of novel architectures for specific applications.¹

Our proposed mechanism of the electrochemical doping of ProDOT(OE₃)-DMP, as presented in Figure 1, will be elaborated through the discussion of the experimental results and the implications thereof. We highlight the significance of these interpretations as they relate to the

optoelectronic properties of electrochemically induced charges and their interaction with neutral excitons, especially in the context of recent literature concerning the excited state dynamics of an acid-doped polyelectrolyte.³⁸ These foundational studies demonstrate the utility and application of *in situ* ultrafast and nonlinear spectroscopy in understanding redox processes in OMIECs.

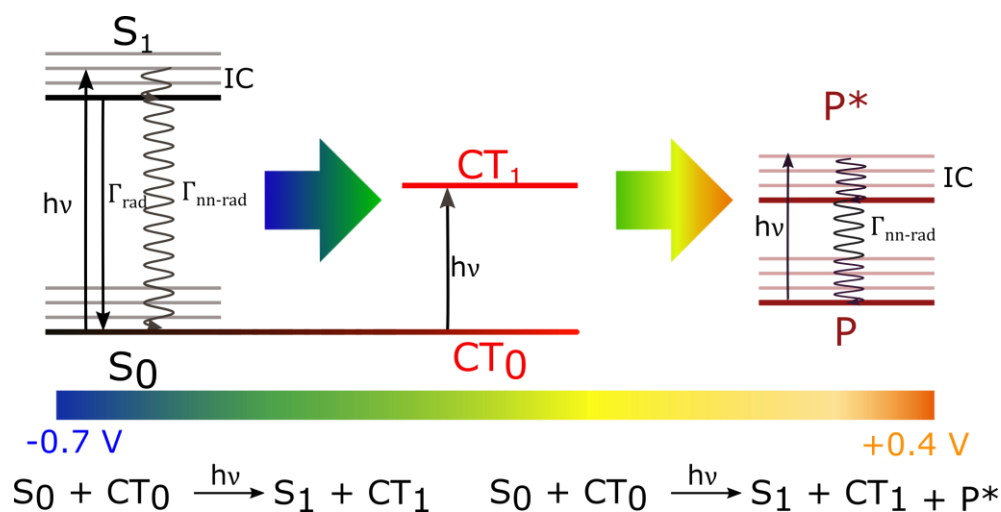


Figure 1: Schematic representation of the electronic states in ProDOT(OE₃)-DMP and their evolution with doping. In the neutral state the primary photoexcitation is constituted by singlet excitons (S₁). A polarized molecular excitation, denoted as a charge transfer state (CT), shares a common ground state with the exciton, appears before substantial doping, and acts as a precursor to polaron (P) formation. Pathways of radiative decay/fluorescence (Γ_{rad}), nonradiative decay (Γ_{nn-rad}), and internal conversion (IC) are labeled.

Experimental Section

Film preparation

The polymer was synthesized via direct (hetero)arylation polymerization by copolymerizing a 3,4-propylenedioxythiophene (ProDOT) comonomer (difunctionalized at the 2-position of the propylene bridge with two oligoether (OE₃) chains) with a 3,3-dimethyl-3,4-dihydro-2H-thieno[3,4-*b*][1,4]dioxepine comonomer (referred to as 2,2-dimethyl-ProDOT or DMP). The resulting polymer is referred to here as ProDOT(OE₃)-DMP. The repeat unit structure is shown in **Figure 2A**. The molecular weight of the polymer was determined to be $M_n = 23$ kDa ($\mathcal{D} = 2.0$) by

GPC in CHCl_3 (40 °C) vs. a polystyrene standard. The synthetic conditions, $^1\text{H-NMR}$, $^{13}\text{C-NMR}$, elemental analysis, and size exclusions chromatograms can be found in previous literature.³¹ Due to the polar oligoether chains and electron-rich backbone, this polymer supports the transport of hydrated/solvated ions and is highly redox active in not only organic electrolytes but also aqueous electrolytes and buffers. Accordingly, films of this material can be fully electrochemically p-doped (complete depletion of the neutral π - π^* optical transition) within the electrochemical window of water. For all *in situ* characterizations, thin films of ProDOT(OE₃)-DMP were spray-cast from room-temperature chloroform solutions (4 mg mL⁻¹) onto ITO/glass slides (8–12 ohm sq⁻¹, Delta Technologies) using an Iwata-Eclipse HP-BC airbrush with argon carrier gas (20 psi). Films with thicknesses ranging from 100–200 nm were obtained. Prior to coating, ITO/glass substrates were cleaned by sonication in sodium dodecyl sulfate in deionized water, deionized water, acetone, and isopropanol, then dried under an argon stream.

X-Ray Scattering

Films of ProDOT(OE₃)-DMP were spray-cast as previously described onto polished silicon wafers. Prior to coating samples, the wafers were cleaned with Piranha etch, then rinsed with distilled water and dried under an argon stream. Grazing incidence wide angle X-ray scattering (GIWAXS) was performed at the Stanford Synchrotron Radiation Lightsource on Beamline 11-3. The sample was irradiated with an incident beam energy of 12.7 keV with a sample distance of 250 mm. Calibration of the scattering profile was performed using a standard sample of LaB₆. More information about GIWAXS data analysis can be found in the Supporting Information (SI Figure S1, SI Table S1).

UV-Vis-NIR Spectroelectrochemistry

Polymer-coated ITO/glass slides were immersed in a standard three-electrode electrochemical cell containing 0.5 M NaCl (biotechnology grade, Amresco) dissolved in ultrapure H₂O (> 18.2 M Ω), a platinum flag counter electrode, and a reference electrode (Ag/AgCl, 3M aqueous NaCl inner solution, BASi, Inc.). Electrolyte pH conditions evaluated using Micro Essential Lab Hydrion Insta-Check pH Test Paper (Ficher Sci.). Before recording UV-Vis-NIR absorbance spectra in the aqueous media, the films were electrochemically conditioned by cycling them between -0.8 and +0.8 V vs. Ag/AgCl for 10 cycles at 50 mV s⁻¹. Potential-dependent linear

absorption spectra were recorded using a Cary 5000 UV-Vis-NIR spectrophotometer (300–1400 nm, 2000 nm s⁻¹) using a quartz cuvette (path length of 1 cm) under ambient conditions. The electrochemical cell was addressed using an EG&G Princeton Applied Research model 273A potentiostat/galvanostat under CorrWare control. For experiments performed in an organic electrolyte, films were prepared in the same manner and incorporated into a three-electrode cell containing a 0.5 M tetrabutylammonium hexafluorophosphate (TBAPF₆, Acros Organics, 98%, purified by recrystallization from hot ethanol) solution in propylene carbonate (PC, Acros organics, 99.5 %, purified using a solvent purification system). The counter electrode was a platinum flag, and the reference electrode was a Ag/Ag⁺ pseudoreference electrode (10 mM AgNO₃ in 0.5 M TBAPF₆ in acetonitrile as the inner solution, $E_{1/2}$ for Fc/Fc⁺ couple = +0.056 V in a 0.5 M TBAPF₆/PC electrolyte solution under ambient conditions. Prior to spectroelectrochemical measurements, films were electrochemically conditioned by cycling them between -0.80 and +0.80 V vs. the Ag/Ag⁺ reference for 10 cycles at 50 mV s⁻¹. Prior to recording all spectra, films were held at the indicated electrochemical conditions for 15–30 seconds until the current transient reached a steady state, indicating an equilibrated oxidation level. Cyclic voltammograms are shown for films spray-cast on ITO/glass in the indicated electrolyte, recorded at 50 mV s⁻¹. The reported voltammograms represent the trace obtained after 5–15 cycles to ensure proper conditioning of the film and reproducible current-voltage characteristics.

Transient Absorption Measurements

Measurements were performed similarly to the linear spectroelectrochemical measurements described above, using the same cuvette, electrolytes, electrode architecture, and film deposition conditions, as shown in **Figure 5A**. Prior to data acquisition, freshly cast films were conditioned by repeated electrochemical cycling as described previously. The electrochemical cell was controlled by a Pine WaveNow Potentiostat and AfterMath software. The transient absorption spectroscopy configuration was commercially acquired (Hera system, Light Conversion) and used a 20 W laser source emitting ~220 fs pulses at 1030 nm with 100 kHz repetition rate (Pharos Model PH1-20-0200-02-10, Light Conversion). An optical parametric amplifier (Orpheus, Light Conversion) was used to generate the pump excitation (2.29 eV, 542 nm). A single-filament white-light continuum (1.18–2.53 eV, 490–1050 nm) was generated by focusing 2 W of the laser light onto a sapphire crystal and used as the probe beam. The probe light transmitted through the sample

was dispersed by an imaging spectrograph (Shamrock 193i, Andor Technology Ltd., U.K.) and detected by a multichannel detector. For the transient spectra recorded as a function of the applied potential, fast scans of the time delay between pump and probe beams were used, where only three delays were acquired. For the dynamic measurements, the time delay ranged from -1 to +10 ps. The pump excitation densities were scaled according to the loss of absorption measured by linear absorption. For the neutral state of the film, the typical fluence used was $0.2 \mu\text{J}/\text{cm}^2$, with a pump beam radius of $\sim 469 \mu\text{m}$ estimated at the $1/e^2$ plane.

Raman Spectroscopy

Resonance Raman spectroscopy was conducted at the Center for Nanophase Materials Sciences at Oak Ridge National Laboratory (Oak Ridge, TN). Measurements were performed *in situ* using a custom-made cell filled with 0.5 M NaCl/H₂O electrolyte, as shown in **Figure 4A**. For sample preparation, films of ProDOT(OE₃)-DMP were spray-cast on ITO/glass working electrodes as described above. The cell was equipped with the same Pt counter electrode and Ag/AgCl reference electrodes identified previously. The polymer-coated working electrode was placed on top of the cell with the polymer-coated side facing downward to contact the electrolyte. The electrochemical cell was controlled by a Pine WaveNow Potentiostat and AfterMath software. The films were electrochemically conditioned in 0.5 M NaCl/H₂O as previously described and allowed to equilibrate for 10–30 seconds at each potential prior to spectral acquisition.

Raman spectra were collected in reflectance mode using inVia Renishaw confocal micro-Raman spectroscopy systems equipped with a 532 nm (2.44 eV) laser and a 758 nm (1.58 eV) laser. Raman spectra were collected while applying the indicated electrochemical potential. For both excitation energies, the excitation lasers were focused through the back of the ITO/glass substrate and onto the underlying polymer film using a 20x objective lens (10% power, exposure time of 10 s, slit opening of 100 μm). For the 1.58 eV excitation, spectra were collected from 100–3000 cm^{-1} using a grating with 1200 lines mm^{-1} . For the 2.44 eV excitation, spectra were collected from 8–2000 cm^{-1} using a grating with 3000 lines mm^{-1} . For each measurement, 6–9 individual acquisitions were performed on different points on the film and averaged. The potential-dependent spectra for multiple films were collected to ensure reproducibility. Representative Raman spectra for each excitation wavelength correspond to data taken with two separate films of comparable

thicknesses. Data analyses (baseline subtraction, cosmic ray removal, and peak fitting) were performed in the Renishaw WiRE 5 software.

Results and Discussion

Microstructural Analysis

The factors governing charge carrier generation, stabilization, and transport in conjugated polymers are complex, often depending on a range of interrelated chemical, electronic, and structural properties. For neat polymer films in the solid state, the macromolecular arrangement of chain segments is known to play a critical role in shaping the optoelectronic properties of the bulk material. To probe the extent of solid-state ordering in the polar conjugated polymer ProDOT(OE₃)-DMP (**Figure 2A**), grazing-incidence wide angle X-ray scattering (GIWAXS) was performed on neat films of the polymer spray-cast from a dilute chloroform solution. The two-dimensional GIWAXS diffraction pattern of ProDOT(OE₃)-DMP (**Figure 2B**) does not suggest the presence of any significant crystalline microstructure or long-range order. The diffracting regions of the film exhibits a slight face-on texture relative to the substrate, as indicated by the notable out-of-plane (q_z) scattering at $q_z = 1.56 \text{ \AA}^{-1}$, which is attributed to scattering from π stacks (labeled (020) in **Figure 2B**, $d_{020} = 4.03 \text{ \AA}$). This feature is superimposed on a broad, amorphous scattering halo, commonly attributed to side chain disorder in alkyl-functionalized polythiophenes.³⁹ Scattering in the q_z direction at $\sim 0.36 \text{ \AA}^{-1}$ is attributed to lamellar stacking (denoted (100) in **Figure 2B**) and corresponds to a distance between adjacent polymer chains of $d_{100} = 17.5 \text{ \AA}$.

The extent of disorder in any given crystallographic direction can be quantified by the paracrystallinity parameter g , which encompasses static fluctuations around a given lattice spacing. Here, we find the g for ProDOT(OE₃)-DMP to be approximately 19% in the $(0k0)$ direction (SI Table S1), which indicates an amorphous film microstructure⁴⁰ and suggests that the intermolecular order associated with π -stacks is not maintained for an appreciable length scale. For comparison, semicrystalline poly(3-hexylthiophene) (P3HT) typically exhibits relatively tighter π -stacking distances ($d_{020} = c. 3.8 \text{ \AA}$)^{28,39,41-43} as well as much lower paracrystallinity in the

($0k0$) direction ($g = 4-7\%$,⁴⁴ depending on molecular weight and processing conditions), indicating significantly less disorder in the system.

The 2D GIWAXS pattern of spray-cast ProDOT(OE₃)-DMP appears to be more consistent short-range, weakly diffracting aggregates. Aggregates, which are distinct from crystallites, are comprised of chain segments that exhibit order in only one dimension (e.g. π -stacking) without participating in complementary structure along crystallographic directions (e.g. lamellar ordering).⁴⁵ For ProDOT(OE₃)-DMP, such aggregation is apparent from the scattering pattern (**Figure 2B**) and corrected line cuts (SI Figure S1), which show weak out-of-plane π - π scattering features that are not complemented by any in-plane lamellar scattering of comparable intensity. Similarly, the relatively out-of-plane lamellar scattering (100) is not complemented by any significant in-plane diffraction from π stacks. We note that the coexistence of scattering from (020) and (100) in the q_z direction suggests the presence of two distinct populations of aggregates. While aggregates are necessary for the formation of bulk crystallites, they may also exist in amorphous matrices. The GIWAXS pattern for ProDOT(OE₃)-DMP suggests that, unlike many widely studied conjugated polymers, this material does not assemble into domains with intermolecular order spanning multiple dimensions. Rather, we see that any ordering of chain segments in ProDOT(OE₃)-DMP is incredibly short range and confined to one dimension only. The lack of multi-dimensional ordering observed in ProDOT(OE₃)-DMP is confirmed by the absence of any prominent thermal transitions in differential scanning calorimetry (SI Figure S2), as reported previously.³¹ Accordingly, X-ray methods are not expected to be particularly powerful tools for developing a structural picture of charged states in this material. We therefore employ *in situ* spectroscopic techniques as more relevant, molecular-level methods for understanding the electrochemical doping of ProDOT(OE₃)-DMP.

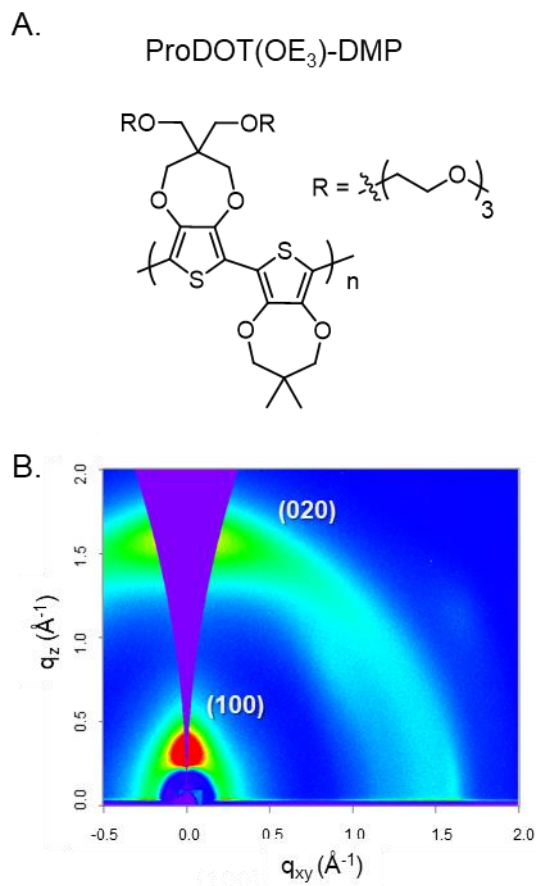


Figure 2: (A) Repeat unit structure of ProDOT(OE₃)-DMP. (B) Two-dimensional GIWAXS diffraction patterns of spray-cast films of ProDOT(OE₃)-DMP on polished silicon, with lamellar (100) and π - π (020) diffraction arcs labeled. Scattering intensity is represented linearly on a blue (low) to red (high) color scale.

Electrochemical Doping of ProDOT(OE₃)-DMP

Next, we focus on the redox behavior of ProDOT(OE₃)-DMP and its potential-dependent linear absorption spectra to understand the doping processes in this material. Electrochemical doping affords a precise and highly controlled means to modify the charge-carrier density (and therefore the oxidation state or doping level) of conjugated polymers. **Figure 3A** shows the cyclic voltammogram (CV) of a spray-cast ProDOT(OE₃)-DMP film in ambient 0.5 M NaCl/H₂O. The potential sweep ranges from -0.8 V vs. Ag/AgCl, where the film is in its charge-neutral state, to

+0.8 V, which is sufficiently high to convert the film to its p-doped form while avoiding concomitant water electrolysis. The current response shows two broad redox processes with an onset of oxidation at approximately -0.21 V. The onset of oxidation depends not only on the ionization energy of the polymer but also on the ability of the polymer to delocalize charge carriers along and across chains, the favorability of electron transfer at the polymer-electrode interface, and the ability of the electrolyte to penetrate the polymer film and allow for facile ion transport in and out of the film. Here, the low oxidation potential of ProDOT(OE₃)-DMP is due to a combination of π electron-donating oxygen atoms in the β -positions of the thiophene ring that increase the overall electron density of the conjugated backbone, as well as the polar side chains that promote the uptake of water and hydrated ions.^{31,32} Following the principal onset of oxidation, an additional oxidation wave is observed, with an onset of approximately +0.2 V and a peak current observed at +0.45 V. In alkyl-functionalized polythiophenes and similar ProDOT-based copolymers, multiple redox peaks have been proposed to arise from the presence of electronically resolved subpopulations that vary in their interchain packing and disorder that give rise to the coexistence of film domains with different potential energy landscapes.^{46–48}

We can utilize the fact that the neutral and charged states of the polymer exhibit characteristic optical absorption bands in the visible and near-infrared (NIR) regions by tracking the potential-dependent optical lineshape with linear spectroscopy. This allows us to monitor the evolution and extent of the electrochemical doping reaction. Linear absorption spectra of a ProDOT(OE₃)-DMP film were recorded in 0.1 V increments from -0.7 V to +0.8 V (**Figure 3B**, plotted in logarithmic scale; see SI Figure S3 for spectra plotted on a linear scale). The absorbance signature of the neutral polymer exciton appears in the 1.8—3.0 eV energy region, showing vibronic features $A_{0,0}$ at ~1.94 eV, $A_{0,1}$ at ~2.13 eV, and a high energy shoulder $A_{0,2}$ at ~2.32 eV. During the initial portion of the potential sweep from -0.7 to -0.2 V (**Figure 3A**, Region I, highlighted in blue), the linear spectra show the neutral exciton as well as two small absorbance features in the NIR at ~1.58 eV and ~1.11 eV. Immediately prior to the onset of oxidation (ca. -0.21 V), these two spectral features in the NIR gain strength, which is accompanied by a moderate bleaching of the neutral ground state. Such features continue to grow immediately after the onset of the doping process between -0.20 V and 0.00 V. In addition, we observe that the feature at ~1.11 eV shows a vibronic structure (SI Figures S3, S8). This electron-vibrational coupling is affected by doping; an empirical baseline

subtraction shows that the vibronic progression associated with this state blue-shifts, and that the peak at the lowest energy loses relative intensity as the doping proceeds (SI Figure S8).

Given that these features exist already in a potential range where there is negligible redox current, they cannot be representative of polarons that form upon oxidative doping.³⁷ Moreover, we observe that the intensity of these transitions in the de-doped polymer depends on subtle changes in the pH of aqueous electrolyte, with more acidic aqueous media giving rise to increased absorbance at 1.58 and 1.11 eV (SI Figure S4). Meanwhile, the choice of processing method (spray vs. blade coating) shows minimal impact on the presence of these sub-gap features (SI Figure S4). To exclude the possibility that these features are associated with any type of protic acid doping process unique to aqueous media, we also evaluated the cyclic voltammetry and potential-dependent spectra for ProDOT(OE₃)-DMP in a polar aprotic electrolyte solution of TBAPF₆ in propylene carbonate (PC). In this environment, such features are still evident and grow in intensity up to 0.2 V *prior to* the onset of oxidation (SI Figure S5). As a structural comparison, we examine the cyclic voltammetry and potential-dependent spectra for two similar ProDOT-based copolymers functionalized with aliphatic and ester side chains that are considerably less polar than the oligoether chains in ProDOT(OE₃)-DMP. When doping these less polar polymers in the polar aprotic medium, we still observe sub-gap such features at ~1.6 eV. However, these features are slightly blue-shifted, are less intense in the de-doped state, and exhibit a slightly different potential dependence when compared to the sub-gap states observed for ProDOT(OE₃)-DMP (SI Figures S6-S7). From these observations, we conclude that the energy and intensity of these sub-gap species are highly dependent on the environment and appear to be stabilized by polarity, be it polarity induced by electrolyte composition or polymer structure. We will expand on the description of these NIR features later in the manuscript, focusing on the transition at 1.58 eV.

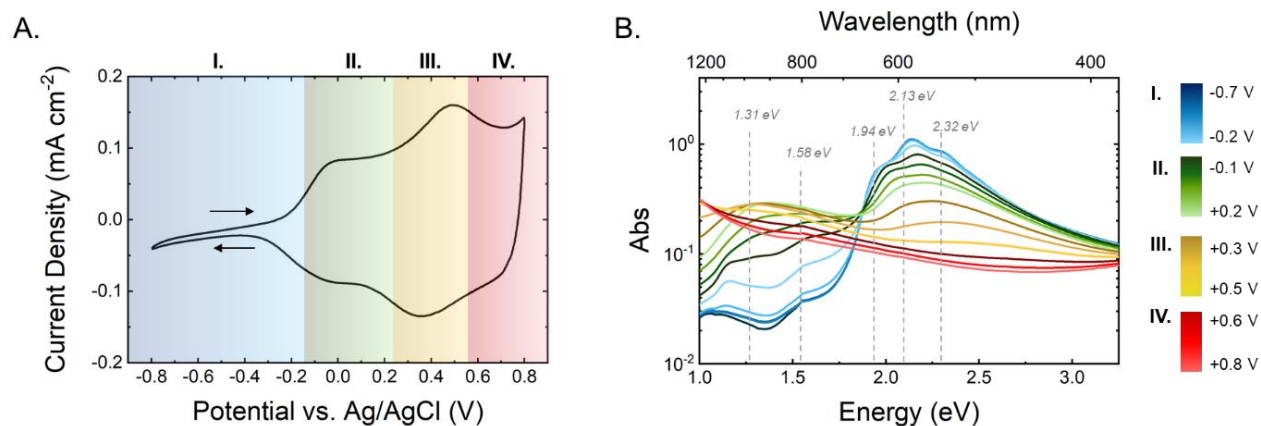


Figure 3: (A) Cyclic voltammogram of ProDOT(OE₃)-DMP on ITO/glass in 0.5 M NaCl/H₂O recorded under ambient conditions at 50 mV s⁻¹ from -0.8 V to +0.7 V vs. Ag/AgCl. Arrows indicate the direction of the potential waveform. (B) Linear absorption spectra of ProDOT(OE₃)-DMP as a function of potential from -0.7 V to +0.8 V in 0.1 V increments. Potential regions I, II, III, IV are indicated in blue, green, yellow, and red gradients, respectively.

After the onset of oxidation of ProDOT(OE₃)-DMP in 0.5 M NaCl/H₂O (-0.21 V), a significant amount of charge is extracted from the film, as indicated by a high redox current between -0.1 and +0.3 V vs. Ag/AgCl (**Figure 3**, Region II, highlighted in green). In this potential range, a new spectral feature emerges in the NIR centered at ~1.31 eV. Throughout this redox process, the new feature at 1.31 eV increases in intensity at the expense of the neutral state absorption through an isosbestic point at 1.88 eV. *In situ* electrochemical conductance measurements suggest that this species is responsible for modulating the bulk electrical conductivity of ProDOT(OE₃)-DMP. The conductance peaks with the maximal absorbance at 1.31 eV (ca. +0.4 V), and the conductance begins to drop at potentials where this spectral feature decreases in intensity (SI Figure S9). Hence, we infer that the charged state absorbing at 1.31 eV corresponds to the primary mobile carriers in the system. Therefore, we attribute the state at 1.31 eV to mobile polarons formed upon electrochemical doping.

For a limited range of potentials (below +0.3 V), the influence of this polaron feature does not obfuscate the structure of the neutral excitonic transition significantly. Therefore, within this potential window, we can evaluate the evolution of the main vibronic features A_{0-0} and A_{0-1} in terms of their energy and relative intensities. First, the energetic difference between the two principal vibronic features ($E_{vib} = E_{0-0} - E_{0-1}$) decreases by a factor of ~ 0.8 , indicating a relaxation of the C=C double bond in the polymer backbone. In addition, the exciton absorbance spectrum blue-shifts considerably throughout the doping process, with the energy of the main A_{0-1} transition increasing by approximately $\sim 63 \pm 6$ meV (SI Figure S10). These results suggest a net disruption of the macromolecular structure and conformation of the neutral chains giving rise to the neutral ground state population. Moreover, the absorbance ratio between the main two vibronic peaks (A_{0-0}/A_{0-1}) increases by a factor of ~ 1.7 between -0.4 and +0.2 V. A similar evolution of vibronic features has been observed for chemically and electrochemically doped semicrystalline P3HT,^{28,49,50} where the lineshape of the UV-Vis absorption spectrum may be used as a diagnostic proxy for studying the nature of HJ aggregation in the film.^{51,52} Thomas et al. attributed the increasing A_{0-0}/A_{0-1} ratio in electrochemically-gated P3HT to the preferential doping of H-aggregates,²⁸ which was later supported by similar studies by Harris et al. that described conformational relaxations associated with heterogeneous electrochemical doping.⁴⁹ Such proposed physical pictures are consistent with previously mentioned work hypothesizing the presence of electrically resolvable chromophore populations with different effective conjugation lengths that are sequentially and selectively p-doped with increasing potential.⁴⁶ In the case of ProDOT(OE₃)-DMP, we consider the additional possibility that the observed changes in the neutral lineshape during the first redox process could also be explained by distortions in chain conformation due to ion intercalation or interactions of excitons with nascent polarons. Even in the case of preferential doping of specific film domains or regions, electrolyte uptake within and around doped regions may influence the conformation of neighboring neutral domains, resulting in notable changes in the absorption lineshape. In other words, chain segments that are not oxidized during this first redox process do not necessarily remain unaffected by electrolyte uptake and charge formation occurring elsewhere in the film.

The second oxidation process in the CV occurs between approximately +0.3 V to +0.5 V (**Figure 3**, Region III, highlighted in yellow). During this process, the neutral state absorption

continues to blue shift and loses its fine structure. One hypothesis to explain this trend would be that the remaining neutral chromophores become mostly uncoupled at higher doping levels and see an increase in inhomogeneous disorder width of the neutral exciton. However, we note that the absorption lineshape and linewidth are influenced both by macromolecular structure and by the environment, on corresponding length scales. Changes in the linewidth represent both changes in static disorder via the inhomogeneous linewidth, as well as dynamic disorder and environmental fluctuations via the homogeneous linewidth. It is therefore impossible to identify the spectral shift and broadening due to disorder from only changes in the linear absorption lineshape. Nonlinear coherent methods would be required to separate homogeneous and inhomogeneous contributions in polymer systems.^{53,54}

This potential window also corresponds to a loss of the polaronic absorption at 1.31 eV and the emergence of another feature further into the NIR with an isosbestic point around 1.15 eV. This low-energy feature is often attributed to bipolarons in the conventional charge carrier model.^{55,56} Above +0.6 V (**Figure 3**, Region IV, highlighted in red), the neutral state absorption is completely depleted, while the 1.31 eV polaron feature continues to decrease in intensity up to the anodic switching potential (+0.8 V). Upon reversing the direction of the potential sweep, the neutral state of the polymer is fully recovered. The amount of charge extracted during the oxidation process (2.2 mC cm⁻²) corresponds to the charge inserted during the reduction process, which provides further evidence that the doping process is repeatable.

The evolution of the linear absorption spectra with electrochemical potential raises several points of interest. *(i)* The lineshape of the neutral state absorption undergoes significant changes even before any substantial redox current is detected and prior to the emergence of polaronic signatures at 1.31 eV. *(ii)* The spectral features at 1.58 eV and 1.11 eV do not share the same potential dependence as the feature at 1.31 eV, which we attribute to polarons. In particular, the state at 1.58 eV gains in strength before the appearance of the polaronic optical signature and before substantial current extraction. Moreover, the state at 1.58 eV appears to be dependent on the properties and polarity of the surrounding environment. This observation suggests the features at 1.58 and 1.11 eV are representative of a precursor species that mediates the electrochemical doping process and precedes polaron formation. As we will show later, the sub-gap state at 1.58 eV is consistent with a CT-like transition. Therefore, we will refer to it as such in later discussion. *(iii)*

The increase in A_{0-0}/A_{0-1} and decrease in E_{vib} of the exciton upon doping suggests that polaron formation influences the electronic configuration of nearby neutral chains. However, linear absorption data alone cannot explain the mechanism of this influence, nor can it describe how the precursor state affects the rest of the doping reaction. Accordingly, we use resonance Raman and transient absorption spectroscopy to gain more insight into the electronic and structural coupling of doped and neutral chromophores and evaluate how this coupling evolves at high doping levels.

Resonance Raman Spectroscopy

To confirm that the NIR state indeed constitutes an electronic transition that is distinct from the neutral exciton and is coupled to a unique molecular conformation, we performed *in situ* resonance Raman spectroscopy utilizing two distinct excitation energies. Raman spectroscopy is a very sensitive probe of the local environment and detects vibrational modes that show a change in the molecular polarizability upon vibration.⁵⁷ In particular, resonance Raman spectroscopy enhances modes coupled to the specific electronic transition in resonance with the laser excitation. This enhancement occurs over orders of magnitude and assists in identifying vibrational modes that are uniquely coupled to distinct electronic transitions. For evaluating films of ProDOT(OE3)-DMP, we employed two excitation energies: one in resonance with the π - π^* electronic transition of neutral chains (2.33 eV) and a second excitation resonant with the sub-gap transition at 1.58 eV.⁵⁸

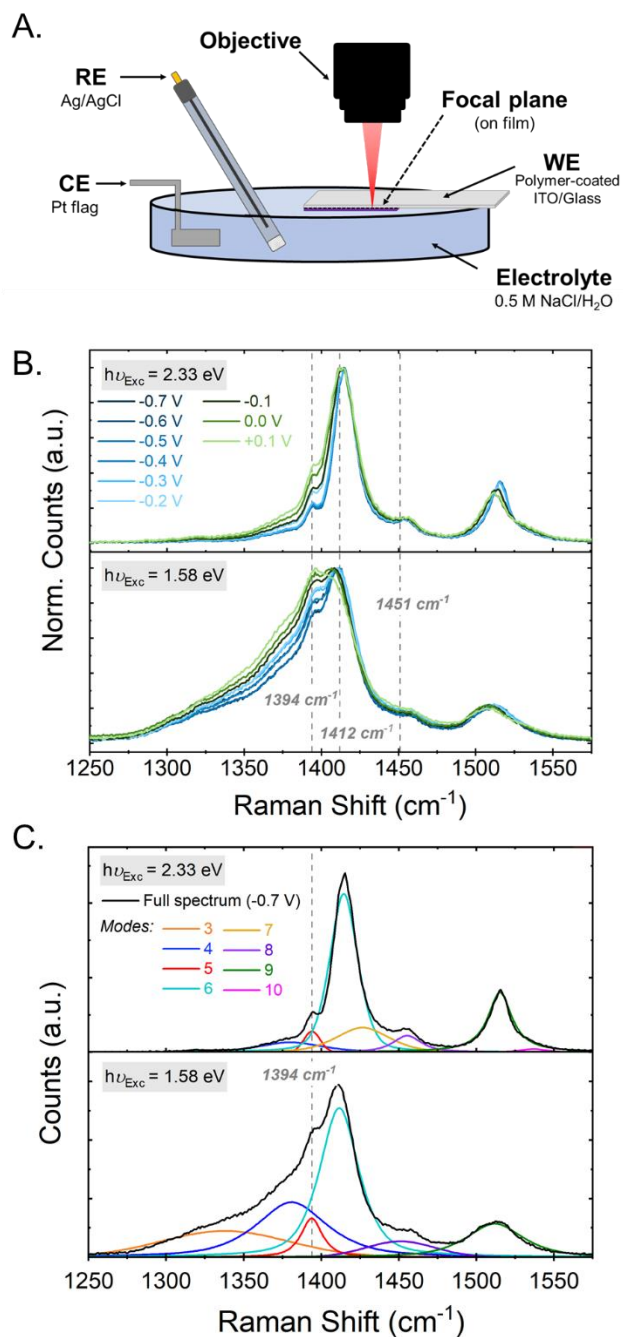


Figure 4: (A) Schematic of cell used for *in situ* Raman measurements performed in an aqueous electrolyte, showing working, reference, and counter electrode (WE, RE, and CE, respectively) and confocal objective. (B) Normalized potential-dependent Raman spectra of ProDOT(OE₃)-DMP using an excitation energy of 2.33 eV, which is resonant with the π - π^* transition of the neutral, de-doped polymer (*top*), and an excitation energy of 1.58 eV, which is resonant with the

sub-gap transition (*bottom*). Spectra were collected *in situ* while biasing the polymer film at the potentials indicated by the color legend. (C) Peak deconvolution for a select region of the Raman spectrum of ProDOT(OE₃)-DMP at -0.7 V for an excitation energy of 2.33 eV (*top*) and an excitation energy of 1.58 eV (*bottom*). Data is shown as a black line with deconvoluted modes superimposed as colored lines and labeled as in Table 1. All data was collected in a 0.5 M NaCl/H₂O electrolyte, and potentials are reported *vs.* Ag/AgCl.

When the polymer is in its charge-neutral state at -0.7 V, the most intense mode for both excitations is at $\sim 1412\text{ cm}^{-1}$ (SI Figure S11), which corresponds to the symmetric stretching of the C $_{\alpha}$ =C $_{\beta}$ double bond in neutral chains.²⁸ When exciting at 1.58 eV, we note a relative enhancement of the modes at $\sim 1338\text{ cm}^{-1}$, $\sim 1381\text{ cm}^{-1}$, and $\sim 1394\text{ cm}^{-1}$ (labeled modes 3, 4, and 5 in **Table 1**). This indicates that these three modes are in resonance with the CT-like transition but not with the neutral electronic transition. This observation establishes that the spectral feature at 1.58 eV indeed corresponds to a distinct electronic transition coupled to a unique molecular geometry. The potential energy surface (PES) of any molecular conformation is determined by the linear combination of all the vibrational modes coupled to it: the PES associated with the CT transition will then be distinct with respect to the PES associated to the π - π^* transition and will have as part of the basis set the modes at frequencies $\sim 1338\text{ cm}^{-1}$, $\sim 1381\text{ cm}^{-1}$, and $\sim 1394\text{ cm}^{-1}$. In other words, the electron distribution and the normal vibrational modes associated with the CT transition are distinct from the ones associated with the neutral configuration. To assign the Raman modes associated with ProDOT(OE₃)-DMP (**Table 1**), we adapt our interpretation from a thorough set of assignments made previously for poly(3,4-ethylenedioxythiophene) (PEDOT), a structurally-similar poly(dioxythiophene).^{59,60} Based on these interpretations, the enhanced vibration at $\sim 1394\text{ cm}^{-1}$ can be assigned to the C $_{\beta}$ -C $_{\beta}$ stretching mode of the thiophene heterocycles. Previous resonance Raman spectroscopy studies on P3HT associated an increase in the intensity ratio of the C $_{\beta}$ -C $_{\beta}$ /C $_{\alpha}$ =C $_{\beta}$ stretching modes with enhanced chain planarity in annealed P3HT samples.^{61,62} Because the C $_{\beta}$ -C $_{\beta}$ mode is enhanced when using an excitation resonant with the CT-like state at 1.58 eV, we hypothesize that this electronic transition is preferentially associated with chromophore segments with enhanced intramolecular ordering and planarity. When evaluating the potential dependence of the Raman spectrum excited with 1.58 eV laser (**Figures 4A and 4B**), we

note that the modes coupled to the CT-state (modes 3, 4, and 5) increase in their relative intensity between -0.6 V and -0.3 V (SI Figures S13, S17): the changes in the Raman modes coupled to the CT-state occur prior to electrochemical doping and before the emergence of the polaronic signature (1.31 eV) at ~0.1 V in the linear spectra (**Figure 2B**). The other modes in the 1.58 eV excitation Raman spectrum do not change considerably before the oxidation onset at ca. -0.2 V (SI Figures S13). Similar changes can be observed in the 2.33 eV excitation spectra (**Figure 4A**), where relative intensity of mode 5 increases prior to the oxidation onset (SI Figures S12, S15). While these changes are less significant than those observed for the 1.58 eV excitation spectra (**Figure 4B**), these observations are consistent with a modulation of the ground state geometry preceding electrochemical doping. As described above, previous characterization of the Raman modes of polythiophenes suggests that the ground state modulation identified here for ProDOT(OE₃)-DMP may be associated with planarization of the thiophene backbone. As the excitation energy resonant with the CT-like state (1.58 eV) provides broader resonance Raman peaks with respect to those obtained by resonantly exciting the neutral segments (2.33 eV excitation energy), we conjecture that the inhomogeneous broadening of these states might be larger than for neutral segments of the polymer chain, suggesting greater microscopic disorder. However, as we cannot separate homogeneous and inhomogeneous broadening with this technique, another possibility is that the intrinsic dephasing rate might be different in the two cases.

Based on these data, the CT-like state at 1.58 eV appears to either passively mediate the electrochemical doping process in ProDOT(OE₃)-DMP or it provides a driving force toward polaron formation by favorably modulating of the PES in a manner that facilitates charge carrier formation. While neither possibility can currently be excluded, as we are unable to distinguish spectator modes from active modes in the Raman spectra, we hypothesize that the CT transition provides a driving force for electrochemical doping. Seeing as the initial stages of doping are also accompanied an increase in the relative intensity of mode 5 (and therefore may also be thought to be associated with increased chain planarity, based on previous literature^{61,62}), we hypothesize that the CT-like transition facilitates doping by enhancing the level of intramolecular order in the polymer, thereby offering greater conformational stabilization to nascent charge carriers.

Table 1: Raman modes of ProDOT(OE₃)-DMP films spray-cast onto ITO/glass substrates and immersed in 0.5 M NaCl/H₂O, applying -0.7 V vs. Ag/AgCl. Mode assignment is based on Raman modes of PEDOT, as reported in the literature.⁵⁹

Mode #	Peak Position for 2.33 eV Excitation (cm⁻¹)	Peak Position for 1.58 eV Excitation (cm⁻¹)	Literature Peak Position for 1.58 eV Excitation (cm⁻¹)	Mode Assignment
1	1103	1085	1097	C-O-C deformation
2^a	1205,1217	1178, 1208	1252	C _α -C _α (inter-ring) stretching
3^b	1320	1337	-	-
4^b	1380	1381	-	
5^b	1394	1394	1370	C _β -C _β stretching
6^b	1414	1412	1414	C _α =C _β (-O) symmetric stretching
7	1426	-	-	-
8^c	1455	1451	1445	C _α =C _β (-O) symmetric stretching of doped segments in PEDOT
9	1515	1512	1516	C _α =C _β (-O) asymmetric stretching
10	1536	-	-	-

^a At -0.7 V, this mode is represented by a single peak in the 2.33 eV excitation spectrum, but appears as two overlapping peaks in the 1.58 eV excitation spectrum. With increasing doping level, the single peak in the 2.33 eV excitation spectrum splits into two distinct signals whose relative intensity varies with potential.

^b Raman modes of ProDOT(OE₃)-DMP that appear to be coupled to the CT-like transition at 1.58 eV.

^c Previous work documenting the Raman modes of PEDOT has ascribed this mode to the C_α=C_β(-O) symmetric stretching of doped segments.⁵⁹ For electrochemically doped ProDOT(OE₃)-DMP, there is not a substantial increase in the intensity of this mode with increasing doping level. In fact, the relative contribution of this mode becomes insignificant at high potentials/doping levels (SI Figures S12, S13). Accordingly, we do not find this mode to be associated with doped chromophores and therefore leave it unassigned for ProDOT(OE₃)-DMP.

For both excitation energies (2.33 and 1.58 eV), Raman mode 6 (~1414 cm⁻¹, corresponding to the C_α=C_β(-O) symmetric stretch) does not exhibit any significant change until after the onset of oxidation at -0.21 V vs. Ag/AgCl. This mode red-shifts during the first redox process (SI Figures S14, S16), representing a relaxation of the C_α=C_β(-O) bond as it adopts a more single-bond character. In P3HT films, a shifting of this mode to lower frequencies has been attributed to greater torsional order and enhanced delocalization of π-electron, which again suggests that the early stages of doping are characterized by enhanced chain planarity.⁶¹⁻⁶³ Similarly, mode 9 (~1515 cm⁻¹, corresponding to the C_α=C_β(-O) asymmetric stretch) also red shifts upon the initial oxidation of the polymer and then blue-shifts after the onset of the second redox process (SI Figure S12, S13). At potentials spanning the second oxidation process (-0.1 to +0.2 V), mode 5 (1394 cm⁻¹, corresponding to the C_β-C_β stretch) first red-shifts, indicating a relaxation of the C_β-C_β bond. After +0.4 V and during the second redox process, this mode shifts to higher energies (SI Figures S14, S16) representing a stiffening of the C_β-C_β bond and a return to its original more double bond character. The blue-shifting of this mode at +0.4 V coincides with the same potential associated with a decrease in intensity of polaronic absorbance at 1.31 eV, the decrease in electrochemical conductance, and the emergence of the low-energy bipolaron signature further into the NIR. Together, these results indicate that during the early stages of doping and charge carrier formation,

the chain segments rearrange themselves, resulting in stiffer inter-ring bonds, relaxed intra-ring bonds, which is indicative of greater chain planarity and torsional order. Meanwhile, the formation of bipolarons stiffens the intra-ring bond. In view of our observations, we can revisit the points raised at the end of our analysis of the linear absorption spectra. Point (i), which concerns changes in the linear absorption occurring prior to any substantial redox current and the appearance of the optical signature of polarons, can be explained by a polarization of the π -electrons in the polymer, which results in a modulation of the energy of the vibrational modes coupled to the neutral electronic transition. This manifests in changes in the resonance Raman spectra, which are similarly observed prior to the oxidation onset of the polymer. The blue-shift of the $C_{\alpha}=C_{\beta}(-O)$ symmetric stretch we observe in the Raman spectra upon doping is also consistent with the reduction in E_{vib} in the linear absorption spectra, most likely due to swelling with electrolyte as counter-ions intercalate into the film to support electronic charges on the polymer backbone. Points (ii) and (iii), which concern the unique potential-dependence of the 1.58 eV state and the apparent influence of charged states on remaining neutral species, are further refined. From the resonance Raman studies, we observe distinct changes in the PES of the polymer for different electrochemical processes. The formation of the polaron feature is associated with a stiffening of the neutral chains and increased overall chain planarity, while further oxidation of the backbone to bipolaronic states stiffens the $C_{\alpha}-C_{\alpha}$ intra-ring bonds as well as the inter-ring bonds. It is interesting to look at the *in situ* conductance in light of this observation (SI Figure S9): it is not until the second oxidation process (ca. +0.3 V) that there is a steep increase in the conductance arising from doping. Mode 2 (the $C_{\alpha}=C_{\beta}(-O)$ symmetric stretch) therefore seems to be associated with the conductivity onset, while the other modes do not show such an association with the conductivity modulation. At higher potentials (above +0.5 V), the conductance then decreases: this corresponds to the potential range where we observe the loss of structure and depletion of the neutral species in the linear absorption data (**Figure 3A**, Regions III and IV).

In Situ Transient Absorption Spectroscopy

As detailed above, the Raman spectroscopy results confirmed the distinct nature of the precursor state at 1.58 eV with respect to the neutral transition, and the potential-dependence of the Raman spectra shed light on the role of this state in the doping process. To address the

formation and evolution of this state at 1.58 eV and its coupling to the neutral excitation, we next performed transient absorption (TA) spectroscopy *in situ* while electrochemically doping ProDOT(OE₃)-DMP in the same electrolyte system, as shown in **Figure 5A**.

TA spectroscopy probes electronic states and their dynamics following photoexcitation. In a typical experiment, a spectrally broad laser pulse is used to probe the changes caused by a second pulse in the optical response of the sample. This second pulse, called the pump beam, is spectrally narrower and temporally delayed with respect to the first pulse, the probe beam. The recorded signal $\Delta T/T = (T_{pumpON} - T_{pumpOFF})/T_{pumpOFF}$ is a function of energy (i.e. inverse wavelength within a factor of hc , where h is the Planck constant and c is the speed of light in vacuum) and corresponds to the relative change in the transmission of the probe induced by the pump, for a fixed time delay. **Figures 5B-C** compare the features of the linear absorption and TA spectra for films of ProDOT(OE₃)-DMP held at the same potential (-0.7 V vs. Ag/AgCl). In **Figure 5B**, the exemplary TA data was collected using a pump energy of 2.29 eV, which matches the absorption of the neutral, de-doped polymer chains.

Depending on the energy of the probe pulse, three different scenarios are possible.⁶⁴ (i) The probe energy is resonant with the electronic transition $S_0 \rightarrow S_n$ (ground state absorption). In this situation, the pump depletes the electronic ground state and the transmission of the sample is greater than for the unperturbed system (positive $\Delta T/T$ signal), giving rise to a photo-bleaching (PB) signature. In general, PB spectrally overlaps with the linear absorption. In the case of ProDOT(OE₃)-DMP, the PB extends from 1.85 eV to ~2.48 eV and bears a similar lineshape as the linear absorption (**Figures 5B-C**). (ii) The probe energy is resonant with transitions $S_n \rightarrow S_{n+1}$. Here, the pump populates these upper excited states; hence $\Delta T/T$ is a negative quantity, which is associated with a photo-induced absorption (PA) signal. In neat, pristine π -conjugated materials, the PA signal is usually assigned to the lowest singlet S_1 excitons.⁶⁴ In **Figure 5C**, PA extends for energies below 1.85 eV, with a notable depression at ~1.58 eV which, as will be elaborated on later, we interpret as a localized PB superimposed on a PA band. (iii) The probe pulse is in resonance with the transition $S_1 \rightarrow S_0$. If this transition is optically allowed, the probe can induce stimulated emission (SE) from the level S_1 to the ground state, which will be perceived as an increase in transmission (positive $\Delta T/T$ signal). The energies of the SE signal typically correspond to those associated with the photoluminescence spectrum. In the case of ProDOT(OE₃)-DMP, this

transition is either forbidden or very weak, so we do not see this signal in our data. By changing the temporal delay between pump and probe beams, it is possible to reconstruct the temporal dynamics of the PB and PA, as shown in **Figure** for time delays of 0.6 ps, 10 ps, 100 ps and 900 ps.

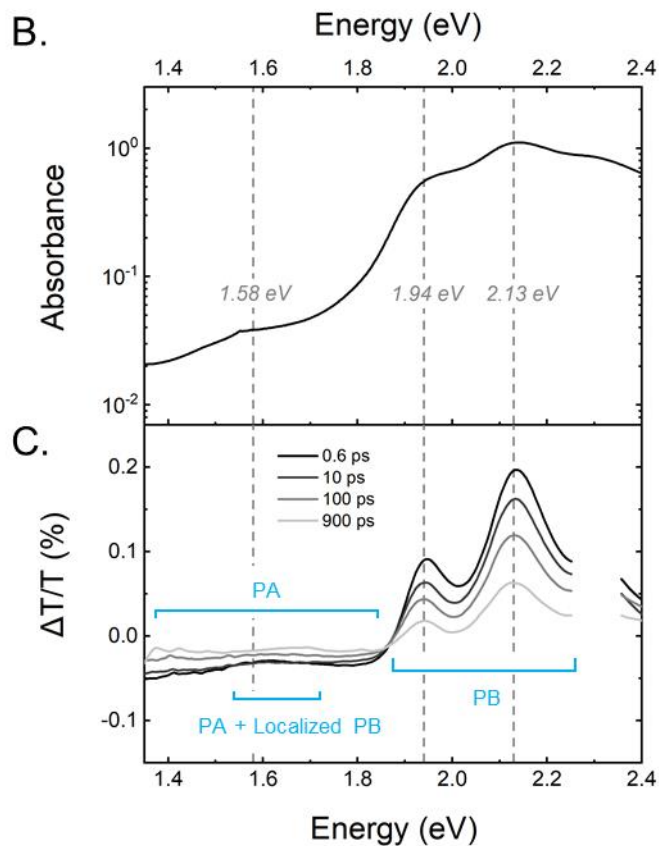
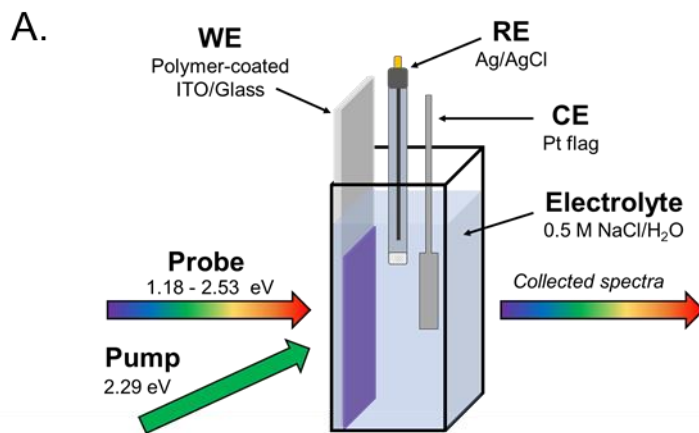


Figure 5: (A) Schematic of the cell used for *in situ* transient absorption measurements performed in an aqueous electrolyte, showing working, reference, and counter electrode (WE, RE, and CE, respectively) and pump/probe pulses. (B) Linear absorbance spectrum of ProDOT(OE₃)-DMP at -0.7 V vs. Ag/AgCl in 0.5 M NaCl/H₂O, highlighting the sub-gap absorbance feature at ~1.58 eV and the vibronic progression of the primary photoexcitation (A_{0-0} at 1.94 eV, A_{0-1} at 2.13 eV). (C) Transient absorption spectra of ProDOT(OE₃)-DMP at -0.7 V for a pump-probe time delays of 0.6 ps, 10 ps, 100 ps, and 900 ps respectively. For clarity, data corresponding to pump energy (2.29 eV) are removed. Regions labeled in blue brackets and text correspond to the photobleach of the primary photoexcitation (PB), the photoinduced absorption (PA), and the localized PB superimposed on the PA occurring at ~1.58 eV.

We will first consider the region corresponding to the PA signal (0.95 eV – 1.85 eV). **Figures 6A-B** compare the linear absorption spectra and TA spectra recorded *in situ* for ProDOT(OE₃)-DMP across a series of electrochemical potentials. The TA spectra presented in **Figure 6B** correspond to a pump-probe delay of 0.6 ps. To highlight the changes in the spectral features, data are shown as absolute values on a logarithmic scale. During the first oxidation process (**Figure 2**, Region II) the PA band arising from the lowest S_1 exciton is quenched (**Figure 6B** and SI Figure S19) and the depression at ~1.58 eV becomes more pronounced with increasing doping level. Because this energy matches the spectral feature at 1.58 eV in the linear absorption data, we interpret this depression as a localized PB signal superimposed to the broader PA band. Such assignment is also illustrated for a single potential (-0.7 V) in **Figure 5B**, for clarity. This localized signal appears soon after the pump interaction (resolution limited ~220 fs), hence it appears to be a common ground state with the neutral exciton (SI Figure 18).

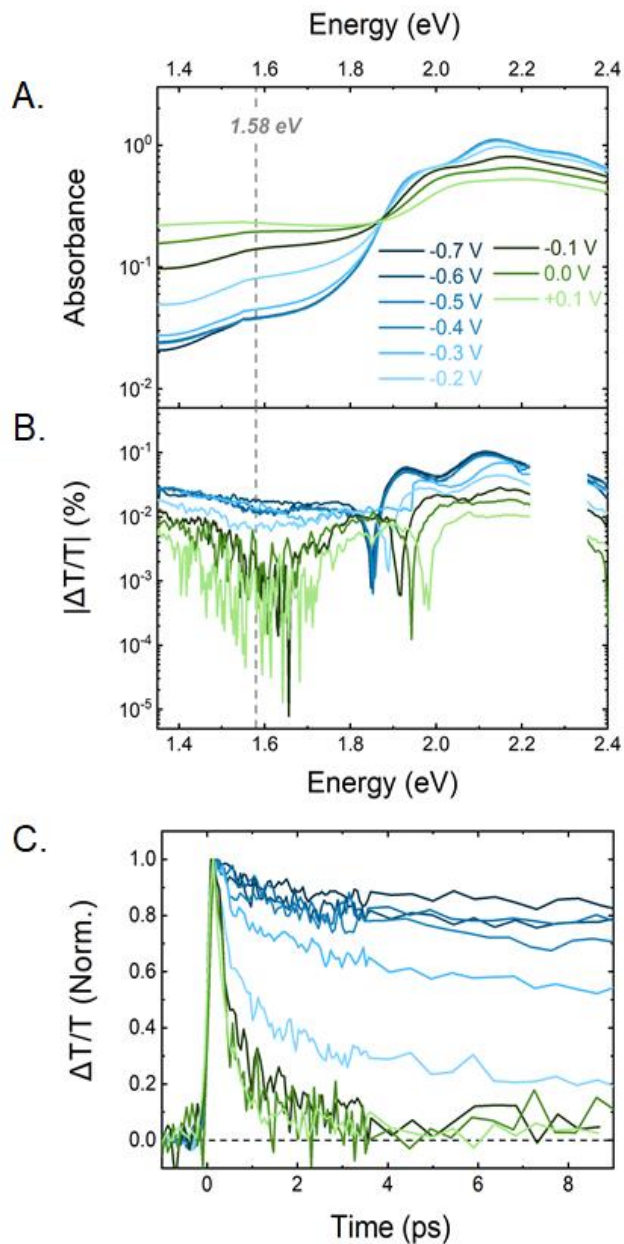


Figure 6: (A) Linear absorption spectra of ProDOT(OE₃)-DMP as a function of potential vs. Ag/AgCl in a 0.5 M NaCl/H₂O electrolyte. From dark blue to light green curves: -0.7 V to +0.1V in even steps of 0.1 V. Note that the absorbance spectra recorded from -0.7 to -0.4 V are similar and therefore overlap. (B) Transient absorption spectra of ProDOT(OE₃)-DMP as a function of potential vs. Ag/AgCl, using identical labeling convention as 4A. Data correspond to a pump-probe delay of 0.6 ps. For clarity, data spanning the pump energy (2.29 eV) are removed. Note that

given the logarithmic scale, the absolute change in transmittance ($|\Delta T/T|$) TA data is plotted. In this convention, negative $\Delta T/T$ are inverted, and the localized PB of the CT-like state is represented by a dip in the PA. (C) Dynamics of the ground state bleach (probe energy = 2.14 eV) from -0.7 V to +0.1 V vs. Ag/AgCl (in steps of 0.1 V) in a 0.5 M NaCl/H₂O electrolyte, using the coloring convention in 4A.

Next, we examine the PB of ProDOT(OE₃)-DMP and its evolution with doping level. We observe that starting from -0.3 V (prior to the current onset in the CV shown in **Figure 2A**), the intensity ratio between the two main peaks identified as A_{0-0} and A_{0-1} (~ 1.92 eV and ~ 2.10 eV) increases. Meanwhile, E_{0-0} blue-shifts as the potential increases, while the spacing between these two peaks E'_{vib} decreases. These changes in the TA line shape are very similar to those we see in the linear absorption data, even though there is a slight discrepancy in the absolute values (SI Figure S10). This observation confirms that the potential dependence of A_{0-0}/A_{0-1} , E_{0-0} , and E'_{vib} is not an artifact arising from interference from the polaron band tail. While the change in the lineshape of the PB is similar to that observed in the linear absorption data, the reduction in its intensity is evidence of a common ground state between the neutral ground state population and the state identified at 1.58 eV, since the former transition undergoes ultrafast quenching ($< \sim 230$ fs, resolution limited). We observe this ultrafast quenching even when progressively increasing the pump excitation densities to compensate for the loss in absorption. For potentials greater than -0.3 V, exciton dynamics become increasingly faster (**Figure 6C**), especially after we reach the potential corresponding to the onset of redox current. We can attribute this behavior to processes of exciton-charge annihilation²⁵ that appears even at the low fluences used here by virtue of the high charge density introduced during electrochemical doping.

Similar results were obtained when the measurements were repeated under an inert atmosphere using a degassed aqueous electrolyte held under an argon blanket during TA measurements (SI Figure S20, S21). Although dynamics are faster in the presence of oxygen, data collected in the degassed environment still show the PB at 1.58 eV superimposed to the PA signature. In addition, an analogous trend is observed in the potential dependence of the PB dynamics. While these results suggest that environmental oxygen does act as a moderate quenching agent with respect to the PB, they also demonstrate that oxygen does not play a substantial role in the formation of the CT-like

state at 1.58 eV, nor does it drastically change the doping process. For comparison, TA experiments were performed using ProDOT(OE₃)-DMP in a polar aprotic electrolyte solution of TBAPF₆ in propylene carbonate (SI Figures S22, S23). While the 1.58 eV state is still present in the de-doped state, the dynamics of the PB recovery are much faster in TBAPF₆/PC and the potential dependence of the quenching is less pronounced when compared to the dynamics observed in NaCl/H₂O. The increased polarity of the aqueous electrolyte and the differences in the charge density evolution for each system may account for the observed differences, as will be explained in the following section.

Discussion

Based on the emergence of the state at 1.58 eV prior to any substantial redox current and polaron formation in the linear absorption, we exclude the possibility that this state represents a polaronic species in the traditional sense. Simultaneously, these results suggest that the state is a precursor to full polaron formation during the doping process. Resonance Raman spectra prove that this state represents an electronic transition that is distinct from the neutral species, as it is coupled to a unique set of vibrational modes. Insights from the transient absorption (TA) spectra show that this species corresponds to a narrow photobleach (PB) that appears contextually to the PB of the neutral transition, indicating that these two states share a common ground state. In sum, these observations lead us to conclude that the precursor state observed at 1.58 eV represents a polarized molecular excitation, i.e. a fundamental molecular excitation which has acquired a charge-transfer character. For this reason, it has been denoted as a charge transfer state (CT-state). This assignment is in agreement with a previous study on the photo-physics of P3HT nanoparticles (NPs) in aqueous buffers.⁶⁵ In the case of P3HT NPs dispersed in ultrapure water, the linear absorption showed a spectral component at low energies (~1.8 eV) that is not present or appreciable in thin films of the same material. This component was enhanced in acidic aqueous buffers (pH values from 3 to 11 were tested). The effect was suppressed for encapsulated NPs, where a lipid-PEG corona was used to shield the NPs from the buffer. TA measurements showed a PB band corresponding to the 1.8 eV feature for NPs in acidic buffers (pH = 3). In this study, these observations were rationalized by assigning the low energy feature to a distribution of CT states due to polarization of π -electrons in the polythiophene backbone. This polarization is

enhanced at low pH, where there are more coulombic interactions between backbone and ionic species, and complexes are formed with electron-withdrawing species that are thought to serve as a precursor for charge generation. Because CT states of this nature are mainly non-radiative, the photoluminescence measurements of these NPs were shown to be insensitive to pH variations.⁶⁵ Here, in the case of electrochemically addressed ProDOT(OE₃)-DMP, the polarization of the excitation could most likely arise from the applied bias and associated ion infiltration into the film. Notably, our TA data collected in an inert atmosphere exclude the possibility that complexation with molecular oxygen plays a central role in the process (SI Figures S20, S21). The physical interpretation of this CT-like photoexcitation is not obvious without detailed electronic structure calculations. The polarization of the molecular excitation could occur along the conjugated polymer backbone as a result of heterogeneous swelling and solvation of the film. A similar explanation was proposed in the study of P3HT NPs, where the origin of the CT state was thought to arise from polarization of “loosely bound” chains experiencing enhanced solvation at the nanoparticle surface.⁶⁵ Alternatively, the CT state in ProDOT(OE₃)-DMP may be inter-chain in origin with polarization occurring between stacked chromophore segments, with polarization occurring between stacked chromophore segments, as was described for helical sexithiophene stacks in work by Glowe et al.⁶⁶ In this study, the authors inferred charge-transfer excitons were localized in adjacent oligomers in the π stack. Such polarization was thought to be stabilized by the polar environment induced by the oligo(ether) end groups. The polar functionality in ProDOT(OE₃)-DMP may contribute to the stabilization of the CT state in an analogous manner. We also consider the possibility that the CT-like state identified in ProDOT(OE₃)-DMP may occur through a combination of both inter- and intramolecular polarization.

The fact that the CT-state is coupled to a unique set of molecular vibrations implies that its electronic configuration and thus its PES are different from those of the neutral ground configuration. In the context of the doping process, this opens up two possible scenarios: on the one hand, the CT-state could play an active part by preparing the PES, favoring the doping of chain and driving the reaction; on the other hand, it could just assume a passive role and therefore not be critical to the doping.

In contrast to the findings of Tsokkou et al., who evaluated acid doping of the donor-acceptor system PCPDTBT,³⁸ we do not observe a broad bleaching feature corresponding to the energy of

the polaron (SI Figure S18). In fact, for ProDOT(OE₃)-DMP, we see no evidence for a shared ground state between the exciton and polaron. Even at +0.1 V vs. Ag/AgCl, when there is a coexistence of both the neutral exciton (~2.13 eV) and the polaron (~1.31 eV), pumping of the exciton does not yield any apparent bleaching of the polaron, as evidenced by the undisturbed PA signal spanning into this range (SI Figures S24, S25). Our evidence suggests that there are subtle differences in the elementary steps governing charge carrier formation in conjugated polymers. These differences, which appear to depend on the mechanism of doping, the extent of doping, as well as the polymer structure, warrant further investigation to better characterize and understand charge carrier formation in conjugated polymers. A notable difference in the work conducted by Tsokkou et al. is the use of an acid to dope a donor-acceptor system which has counter-ions fixed to the side chains (a conjugated polyelectrolyte),³⁸ whereas the work presented here employs electrochemical doping to address an all-donor polymer where the permeation of exogenous counter-ions are necessary to stabilize charge carriers induced by the electrochemical potential. Aside from mimicking the conditions of the targeted application, electrochemical doping configurations offer the unique advantage of precise control of the oxidation state of the polymer in a manner that cannot be easily achieved via redox chemical or acid doping. This feature is of fundamental importance for identifying and understanding the factors that determine the evolution between the ground state and the CT intermediate. Presently, the factors governing the formation of shared ground state complexes are unclear, although differences in the frontier orbital configuration (donor-acceptor or all donor), doping mechanism (chemical, acid, or electrochemical), the presence of tethered counter-ions, film microstructure, and side chain polarity (polar, apolar, or ionic) may play a substantial role.

What seems to be common to both the acid-doped PCPDTBT³⁸ and our analysis of electrochemically doped ProDOT(OE₃)-DMP is the effect of charge concentration on the dynamics of the PB. In both cases, doping is associated with a quenching of the PB, which is to be expected given the general phenomenon of exciton-charge annihilation.⁶⁷ Tsokkou et al. reported dynamics spanning ~1 ps for the fully p-doped polyelectrolyte.³⁸

For ProDOT(OE₃)-DMP, we also note that the excited state dynamics are sensitive to the nature of the electrolyte. In a polar aprotic electrolyte (0.5 M TBAPF₆ in propylene carbonate), the recovery of the ground state bleach is notably reduced, even for similar oxidation states or doping

levels (SI Figure S20). There are several possible origins for this behavior. The higher dielectric constant of the aqueous medium could increase the coulombic screening of the excited state, thereby extending its lifetime. In addition, the de-doped films may also exhibit different degrees of electrolyte trapping after multiple doping/de-doping cycles. In other words, while both films may be in the same oxidation state, the surrounding ionic milieu may solvate each polymer to different extents. Given the known propensity of oligo(ether) chains to chelate small alkali metal ions such as sodium,⁶⁸⁻⁷¹ we expected the electrochemical cycling of ProDOT(OE₃)-DMP in the aqueous sodium chloride electrolyte to potentially result in a greater enrichment of ions in the de-doped film, whereas fewer ions may be trapped when a polar aprotic electrolyte system comprising large polarizable ions (TBAPF₆ in PC). Such dissimilarities in film solvation and ion content could both disrupt the ordering of chromophores and also change the overall dielectric quality of the environment, thereby influencing the excited state dynamics. We note that the absorbance profiles of de-doped ProDOT(OE₃)-DMP appears to have distinct lineshapes depending on the surrounding electrolyte, with the film immersed in NaCl/H₂O exhibiting an enhanced high-energy absorbance feature (SI Figure S26). These are suggestive of subtle changes in the molecular-level arrangement and coupling of chromophores, which may also influence the stabilization and quenching of photoexcitations.

Not only are the dynamics faster in the polar aprotic electrolyte, but the potential-dependence of the recovery dynamics is also weaker when compared to the potential dependence observed in the aqueous medium (SI Figure S23). As the charge densities at the early stages of doping is similar for regardless of the electrolyte (SI Figure S27), these differences cannot be ascribed to different doping densities. One possibility is that the lower dielectric constant of the organic electrolyte provides a reduced Coulomb screening (in the Langevin picture), hence making the interaction between excitons and charges stronger, quenching exciton population in such an effective way in the organic electrolyte that the effect of increasing charge density is reduced. From these findings, we emphasize that the photophysical properties of OMIECs are highly dependent on not only the doping level, but also on the nature and composition of the surrounding electrolyte. The presence of unique polymer-electrolyte interactions give rise to notable differences in optical and electrochemical observables, as shown here for ProDOT(OE₃)-DMP. This suggests a particularly strong coupling of the molecular properties to the surrounding environment and highlights a need

to explore the wide space of environmental parameters (protic and aprotic systems, ion size and structure, polarity, hydrogen bonding, pH, etc.) to develop complete descriptions of photophysical processes in OMIECs.

Conclusions

Ionic-electronic coupling is a fundamental process in OIMECs, and together with ion and electron transport determines the efficiency and functionality of a wide range of materials and devices. In this work we have probed the electrochemical doping mechanism of a redox active ProDOT-based polymer by monitoring linear-, transient-, and Raman spectroscopic variables *in situ* as a function of electrochemical potential. Our findings provide significant insight into the doping mechanism and evaluate the effects of the surrounding environment and intrinsic polymer structure on the doping process. A key finding of our work is the identification of a CT-state (at 1.58 eV) that precedes polaron formation and that shares a common ground state with the neutral transition but not with the electrochemically generated polarons. Using a combination of transient absorption and resonance Raman spectroscopy, we are able to show that the CT-state is stabilized by polarity, be it polarity induced by the electrolyte composition or the polymer structure. In addition, this CT-state is responsible for the quenching of the light-induced exciton population, *via* an ultrafast exciton-charge annihilation mechanism.

Our resonance Raman results showed that polaron formation occurring in the early stages of the doping process also influence nearby neutral chains. In doped segments, polaron formation results in a modification of the chain conformation to a more planar structure with enhanced torsional order and relaxed intra-ring bonds. This seems to be beneficial for charge transport, as we observe a simultaneous increase in electrochemical conductance upon polaron formation. Our TA data confirmed that the electrochemically generated polarons do not share a common ground state with neutral state transition, contrary to what has been shown for an acid-doped donor-acceptor polymers. At high doping levels we observed bipolaron formation accompanied by a stiffening of both the inter- and intra-ring double bonds and a subsequent decrease in electrochemical conductance.

From the TA results, we also conclude that the excited state dynamics are sensitive to the surrounding environment, where substantially faster dynamics were observed in an aprotic organic

electrolyte than in an aqueous electrolyte. By comparing polymers with the same backbone but different substituents in aqueous and organic electrolyte systems, we obtained evidence that subtle changes in molecular arrangement and chromophore coupling influences both the quenching and stabilization of photoexcitations. This highlights the importance and uniqueness of polymer-electrolyte interactions on the photophysical properties of OMIECs and showcases the dependence of these properties on the nature and composition of the surrounding electrolyte.

Across the OMIEC literature, a broad suite of active materials has been reported with varying extents of solid-state order and crystallinity. While a number of OMIECs exhibit substantial solid-state order (especially those used in OECT applications^{35,72,73}), there also exist many high-performing active materials that show little to no long-range order.⁷⁴⁻⁷⁶ In fact, studies have found that greater crystallinity and/or long-range ordering of the π system is not de facto beneficial for performance, especially considering how disordered systems tend to facilitate ion transport.^{30,36,77,78} In addition, OMIECs prepared by electropolymerization (which is a commonly used route, especially for electrochemical supercapacitors⁷⁹) tend to be very disordered. With these examples in mind, we believe that understanding the properties of disordered systems (like ProDOT(OE₃)-DMP) is crucial for the development of the field. Furthermore, because disordered systems tend to be more difficult to characterize, our work also showcases useful handles for studying molecular-level processes in OMIECs lacking long-range order.

Through these efforts, we identify an experimental toolset that is critical for understanding the energy states involved in many electrochemical devices, including (but not limited to) bioelectronic technologies. These pointers can be effectively used to determine how chemical and morphological architectures can be engineered to enhance, suppress, or mediate the states involved in the doping mechanism and mediating the working mechanism of electrochemical doping in conjugated polymers.

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Conflict of Interest

The authors declare the following competing financial interest(s): Electrochromic polymer technology developed at the Georgia Institute of Technology has been licensed to NXN Licensing. JRR and AMÖ serve as a consultant to NXN Licensing.

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Graphical TOC Entry

